

Balancing Equations Answers Key

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The first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation.

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Balancing Equations Worksheet and Key 1. Answer the following questions about the chemical equation shown below: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ a) What are the reactants? H_2 and O_2 b) What is the product? H_2O c) What is the number "2" in front of the H_2 (and H_2O) called? Coefficient d) Is the reaction balanced?

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Balancing Equations: Answers to Practice Problems. 1. Balanced equations. (Coefficients equal to one (1) do not need to be shown in your answers). (a) $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$. (b) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. (c) $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2(\text{SO}_4)_3 + 6\text{HBr}$ (d) $1\text{C}_4\text{H}_6\text{O}_3 + 1\text{H}_2\text{O} \rightarrow 2\text{C}_2\text{H}_4\text{O}_2$.

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Balancing Equations: Practice Problems

Anne Marie Helmenstine, Ph.D. Updated January 04, 2019. A balanced chemical equation gives the number and type of atoms participating in a reaction, the reactants, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow.

How to Balance Equations - Printable Worksheets

Being able to balance chemical equations is a key chemistry skill. Here are the steps involved in balancing chemical equations, plus a worked example. Remember, to balance an equation, you change the coefficients, not the subscripts in the formulas. Once you have balanced one element, do...

Balancing Chemical Equations Phet Answer Key

Tips for Balancing Equations . When balancing equations, remember chemical reactions must satisfy conservation of mass. Check your work to make certain you have the same number and type of atoms on the reactants side as on the products side. A coefficient (number in front of a chemical) is multiplied by all the atoms in that chemical.

Balancing Equations Chemistry Test Questions

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Balancing Chemical Equations – Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$ 7) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ 8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 9) $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$ 10) $1 \text{ FeCl}_3 + 3 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_3 + 3 \text{ NaCl}$ 11) $4 \text{ P} + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$

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