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$K = \frac{P_i}{P_i^0} = \frac{f_i}{f_i^0}$ Use of the Lewis and Randall rule gives:
 $f_i = P_i \phi_i$; $\phi_i = \frac{f_i}{P_i}$; $\phi_i = 1$ atm.

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POWER LAW BASICS | 6 2 POWER LAW BASICS 2.1 ARRHENIUS EQUATION The main assumption behind the Arrhenius expression is that $k = A e^{-E_a/RT}$. This is an approximation, but it works quite well. The rate coefficient is the term that is a function of temperature but may also depend

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Chemical Equilibrium $dG = -SdT + VdP + \sum_{j=1}^n \mu_j dn_j$ μ_j : chemical potential for species j. $G = \sum_{j=1}^n \mu_j n_j$ $T, P = \text{constant}$ $\mu_j = G_j + RT \ln a_j$, $a_j = f_j / f_j^\circ$ $K_i = \prod_{j=1}^n a_j^{\nu_j}$ $G_i^\circ = -RT \ln K_i$ Standard state: pure species j at 1 atm and system temperature.

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Description. Chemical kinetics is the study of the rates of chemical reactions; if changes in conditions impact the speed of a reaction, we can better understand what caused the reaction. In this course, you will learn to understand complicated data sets and analysis techniques for measuring and understanding these rate changes.

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