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Matter : Solid Liquid Gas

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Nature of Matter - Chapter

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11.2 Stoichiometric

Calculations 11.3 Limiting

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Chemistry: Matter and Change

- Chapter 11. STUDY. PLAY.

Mole. Commonly abbreviated

mol, SI base unit used to

measure the amount of a

substance. Avogadro's

Number. 6.02×10^{23} Number

of representative particles

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in a mole of anything. Molar
mass. Mass in grams of one
moles of any pure substance.

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The Stoichiometry chapter of

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11.1 Describing Chemical
Reactions: Introduction to
Chemical Equations: Chemical
reactions involve changing
substances (reactants change

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to one or more products).

Chemists use a chemical equation to convey as much information as possible about what happens in a chemical reaction.

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MANUAL Section 11.1 Answers

Assessment page 372 5.

Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the

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balanced equations indicate
the molar relationship
between each pair of

Stoichiometry Stoichiometry
Chapter 11, Problem 14 is
Solved. The answer to
"Challenge Titanium is a

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transition metal used in many alloys because it is extremely strong and lightweight. Titanium tetrachloride ($TiCl_4$) is extracted from titanium oxide (TiO_2) using chlorine and coke (carbon).

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amounts of reactants used
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formed in a chemical

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MASTERY The Elements Section

7.1 Properties of s-Block

Elements In your textbook,

read about the

representative elements and

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hydrogen. Answer the
following questions. 1.

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