

## Organic Chemistry Hydrocarbons Study Guide Answers

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Propane, butane, and isobutane are all hydrocarbons with only single covalent bonds between carbon atoms. These hydrocarbons that lack double bonds, triple bonds, or ring structures make up the class called alkanes. See Table 1. As the number of carbon atoms increases, so does the number of ways that they can be connected to form different isomers. You should realize that isomers are defined by the pattern of bonding between the carbons.

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In Chapter 1 "Organic Chemistry Review/Hydrocarbons" through Chapter 5 "Amines and Amides", we survey organic chemistry by dividing its compounds into families based on functional groups. We begin with the simplest members of a family and then move on to molecules that are organic in the original sense—that is, they are made by and found in living organisms.

~~Chapter 1 Organic Chemistry Review / Hydrocarbons CHE ...~~

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A more appropriate definition of organic chemistry is: The chemistry of hydrocarbons (compounds of carbon and hydrogen only) and the derivatives of hydrocarbons, (compounds containing C, H, and O, N, S, P, halogens etc). In some organic compounds, the metals such as Na, K, Ca and many others can be found.

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STUDY GUIDE In your textbook, read about organic chemistry, hydrocarbons, and straight-chain alkanes. Use each of the terms below just once to complete the passage. homologous series straight-chain alkanes If all of the carbon atoms are

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A SELF-STUDY GUIDE TO THE PRINCIPLES OF ORGANIC CHEMISTRY The Hydrocarbons chapter of this Glencoe Chemistry - Matter and Change companion course helps students learn the essential chemistry lessons of hydrocarbon molecules. Study Guide for Exam Questions - arri.org

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Organic chemistry and biology overlap often, with one of the best examples being macromolecules. Macromolecules are very large hydrocarbons containing functional groups and additional elements, such as nitrogen and phosphorus.

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Acs Organic Chemistry Study Guide - 08/2020 Organic chemistry is the study of the structure, properties, composition, reactions, and preparation of carbon-containing compounds, which include not only hydrocarbons but also compounds with any number of other elements, including hydrogen (most compounds contain at least one carbon – hydrogen bond), nitrogen, oxygen, halogens, phosphorus, silicon, and sulfur.

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1. Inorganic vs. Organic Chemistry. The study of chemicals is a very specific branch of science, but chemistry can be divided even further into separate fields, two of which are organic and inorganic.

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Alkenes are also simple hydrocarbons, following the formula  $C_n H_{2n}$ . Alkenes differ from alkanes in that they contain at least one double bond. They can also contain many double bonds, but all they need is one to qualify as an alkene. Naming alkenes follows the same rules as alkanes. However, we need to specify where the double bond is placed, and, similar to how side-chains are named, we want the double bond to be in between the two lowest carbons.

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ORGANIC CHEMISTRY II OBJECTIVES FOR FINAL EXAM 1. If given names, identify structures, or if given structures, identify name of each of the following types of compounds: a. Esters b. Amines c. Aromatic compounds including polynuclear aromatic hydrocarbons. 2. Identify structures that are products of the following reactions: a.

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